

MARK SCHEME for the May/June 2014 series

9701 CHEMISTRY

9701/35

Paper 3 (Advanced Practical Skills 1),
maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the May/June 2014 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.

Page 2	Mark Scheme	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2014	9701	35

Question	Sections	Indicative material	Mark	Total
1 (a)	MMO Collection	I Burette readings and correct volume given for dilution and volume between 24.00 and 26.00 cm ³	1	
	PDO Layout	II Initial and final readings and titre value given for rough titre and initial and final readings for two (or more) accurate titrations (<i>minimum of 2 × 2 box</i>)	1	
	PDO Recording	III Appropriate headings and units for accurate titration. and volume FA 3 added recorded for each accurate titre. <i>Headings should match readings.</i> <ul style="list-style-type: none"> initial/start (burette) reading/volume final/end (burette) reading/volume titre or volume/FA 1 used/added (<i>but not “difference”</i>) unit: /cm³ or (cm³) or in cm³ or cm³ for each entry 	1	
	MMO Decisions	IV All accurate burette readings recorded to 0.05 cm ³ . <i>The need to record to 0.05 applies only to the burette readings and not to the recorded titres. Do not award this mark if:</i> <ul style="list-style-type: none"> 50(.00) is used as an initial burette reading more than one final burette reading is 50.(00) V Has two uncorrected, accurate titres within 0.1 cm ³ <i>Do not consider the rough even if ticked. Do not award this mark if having performed two titres within 0.1 cm³ a further titration is performed which is more than 0.10 cm³ from the closer of the initial two titres, unless a further titration, within 0.1 cm³ of any other titration has also been carried out. Do not award the mark if any ‘accurate’ burette readings (apart from initial 0) are given to zero dp.</i>	1	
<p>Round any burette readings to the nearest 0.05 cm³. Check and correct subtractions in the dilution and titre tables. Examiner then selects the “best” titre using the hierarchy: two identical; titres within 0.05 cm³; titres within 0.1 cm³; etc. Candidate scaled titre = Cand vol diluted × Cand mean titre/Sup vol diluted Examiner compares candidate scaled titre with Supervisor’s titre.</p>				
(a)	MMO Quality	VI, VII and VIII Award VI, VII and VIII for $\delta \leq 0.20 \text{ cm}^3$ Award VI and VII for $0.20 \text{ cm}^3 < \delta \leq 0.40 \text{ cm}^3$ Award VI only for a difference of $0.40 < \delta \leq 0.60 \text{ cm}^3$ <i>If the “best” titres are $\geq 0.50 \text{ cm}^3$ apart cancel one of the Q marks.</i>	3	[8]

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Question	Sections	Indicative material	Mark	Total
(b)	ACE Interpretation	<p>Candidate must average two (or more) titres that are all within 0.20 cm³. Working must be shown or ticks must be put next to the two (or more) accurate readings selected.</p> <p><i>The mean should normally be quoted to 2 dp rounded to the nearest 0.01.</i></p> <p><i>Two special cases where the mean may not be to 2 dp: allow mean to 3 dp only for 0.025 or 0.075 e.g. 26.325; allow mean to 1 dp if all accurate burette readings were given to 1 dp and the mean is exactly correct. e.g. 26.0 and 26.2 = 26.1 is correct but 26.0 and 26.1 = 26.1 is incorrect.</i></p> <p><i>Note: the candidate's mean will sometimes be marked as correct even if it is different from the mean calculated by the Examiner for the purpose of assessing accuracy.</i></p>	1	[1]
(c)	ACE Interpretation PDO Display	<p>I Correctly evaluates $\frac{0.500 \times \text{vol diluted}}{250}$ in (i)</p> <p>II Correctly calculates (i) $\times \frac{(b)}{1000}$ in (ii)</p> <p>III Correctly calculates (ii) $\div 2$ in (iii)</p> <p>IV Correctly calculates (iii) $\times \frac{1000}{25}$ in (iv)</p> <p>V All answers given to 3 or 4 sf <i>minimum 3 steps attempted to access this mark</i></p>	1 1	[5]
(d)	ACE Interpretation Conclusion	<p>(i) smallest = 24.70, largest = 24.90</p> <p>(ii) lower conc FA 3 leads to greater titre therefore greater conc. of Ca(OH)₂</p>	1 1	[2]
(e)	ACE Conclusion	limewater reacts with carbon dioxide in air	1	[1]
Qn 1	Total			[17]

Page 4	Mark Scheme	Syllabus	Paper
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Question	Sections	Indicative material	Mark	Total
2 (a)	PDO Layout	I Headings with units and three balance readings and two correctly calculated masses of solid are tabulated clearly. Mass /g, (g), in g, in grams <i>If units are omitted from the headings then they must appear next to each entry in the table.</i>	1	
	MMO Collection	II White solid turns yellow/green-yellow on heating and white/becomes paler on cooling	1	
On Supervisor script calculate mass FA 4 ÷ mass loss to 1 dp. Calculate FA 4 ÷ mass loss for the candidate to 1 dp and calculate difference from Supervisor.				
(a)	MMO Quality	III and IV Award III and IV if $\delta \leq 0.5$.	1	[4]
		Award IV only if $0.5 < \delta \leq 1.0$.	1	
(b)	ACE Interpretation	I Correctly calculates mass loss from results in (a) in (i)	1	[4]
	PDO Display	II Correctly calculates M_r : $\text{ZnCO}_3 = 125.4$ and $\text{CO}_2 = 44(.0)$ in (ii)	1	
		III Uses $\frac{(i) \times 125.4}{44}$ in (ii) (allow ecf)	1	
		IV Uses $\frac{(ii) \times 100}{\text{correct mass FA 4}}$ in (iii)	1	
(c)	ACE Improvement	Heat to constant mass/ use larger mass	1	[2]
	Interpretation	All decomposes/all CO_2 given off smaller % error	1	
Qn 2	Total		[10]	

Page 5	Mark Scheme	Syllabus	Paper
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Question	Sections	Indicative material	Mark	Total								
FA 5 is $ZnCO_3(s) + Al_2(SO_4)_3(s)$ FA 7 is $NH_4Cl(s) + NaHCO_3(s)$												
3 (a)	(i)	MMO Collection	I Effervescence (and colourless solution)	1	[7]							
	(ii)		II White ppt with NaOH soluble in excess	1								
	(iii)		III White ppt with NH_3 insoluble/partially soluble in excess	1								
	(iv)		IV No reaction/no change/no ppt	1								
	(v)		V White ppt (if additional HCl/HNO ₃ added then ppt must remain)	1								
	(vi)	ACE Conclusion	VI Al^{3+} VII SO_4^{2-} as white ppt with Ba^{2+}	1								
3 (b)	(i)	MMO Collection	I Condensation or sublimation (<i>allow misty/white fumes</i>) (on gentle heating) and (gas/ NH_3) turns (damp) red litmus blue	1	[6]							
	(ii)		II White/off-white/cream ppt soluble in NH_3	1								
	(iii)		III No reaction or white ppt in 1 st box (allow no ppt) and effervescence in 2 nd box	1								
	(iv)	MMO Decisions ACE Conclusion	IV Uses limewater in either 3 rd test or in (a)(i) V and VI Identifies NH_4^+ , Cl^- and CO_3^{2-} or SO_3^{2-} with appropriate evidence. V only Identifies all three ions without evidence or identifies two ions with evidence.	1 1								
			<table border="1"> <thead> <tr> <th>ion</th> <th>minimum evidence</th> </tr> </thead> <tbody> <tr> <td>NH_4^+</td> <td>gas turning litmus blue/NH_3 (provided correct litmus result in obs)</td> </tr> <tr> <td>Cl^-</td> <td>white ppt with $AgNO_3$</td> </tr> <tr> <td>CO_3^{2-} SO_3^{2-}</td> <td>(rapid) effervescence with acid white ppt with Ba^{2+} and effervescence/ppt dissolves in acid</td> </tr> </tbody> </table>	ion	minimum evidence	NH_4^+	gas turning litmus blue/ NH_3 (provided correct litmus result in obs)	Cl^-	white ppt with $AgNO_3$	CO_3^{2-} SO_3^{2-}	(rapid) effervescence with acid white ppt with Ba^{2+} and effervescence/ppt dissolves in acid	
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CO_3^{2-} SO_3^{2-}	(rapid) effervescence with acid white ppt with Ba^{2+} and effervescence/ppt dissolves in acid											
Qn 3	Total			[13]								