

# Heat Energy Changes in Chemical Reactions

## Question Paper

Level	Edexcel
Subject	Chemistry
Exam Board	GCSE(9-1)
Topic	Rates of Reaction & Energy Changes
Sub Topic	Heat Energy Changes in Chemical Reactions
Booklet	Question Paper

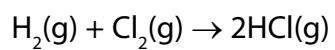
**Time Allowed:** 51 minutes

**Score:** /42

**Percentage:** /100



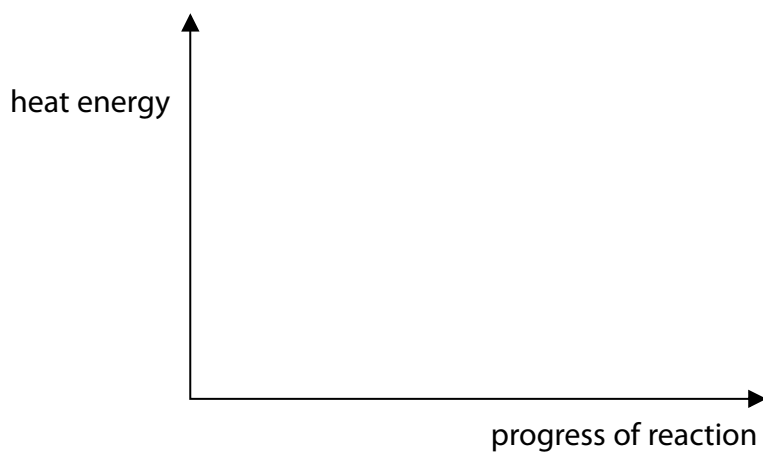
(c) Hydrogen reacts with chlorine to form hydrogen chloride.



The reaction is exothermic.

Draw and label the reaction profile diagram for this reaction, identifying the activation energy.

(3)

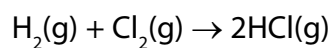


(d) The energies of some bonds are shown in Figure 13.

bond	energy of bond / kJ mol <sup>-1</sup>
H—H	436
Cl—Cl	243
H—Cl	432

Figure 13

Hydrogen reacts with chlorine to form hydrogen chloride.



Calculate the energy change, in kJ mol<sup>-1</sup>, for the reaction of 1 mol of hydrogen gas, H<sub>2</sub>, with 1 mol of chlorine gas, Cl<sub>2</sub>, to form 2 mol of hydrogen chloride gas, HCl.

(4)

energy change = ..... kJ mol<sup>-1</sup>

**(Total for Question 1 = 12 marks)**

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- 2 (a) An experiment is carried out to measure the temperature change when solid ammonium chloride is dissolved in water.

initial temperature of water = 19 °C  
final temperature of solution = 15 °C

Explain what the temperature readings show about the type of heat change occurring when ammonium chloride dissolves in water.

(2)

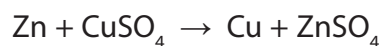
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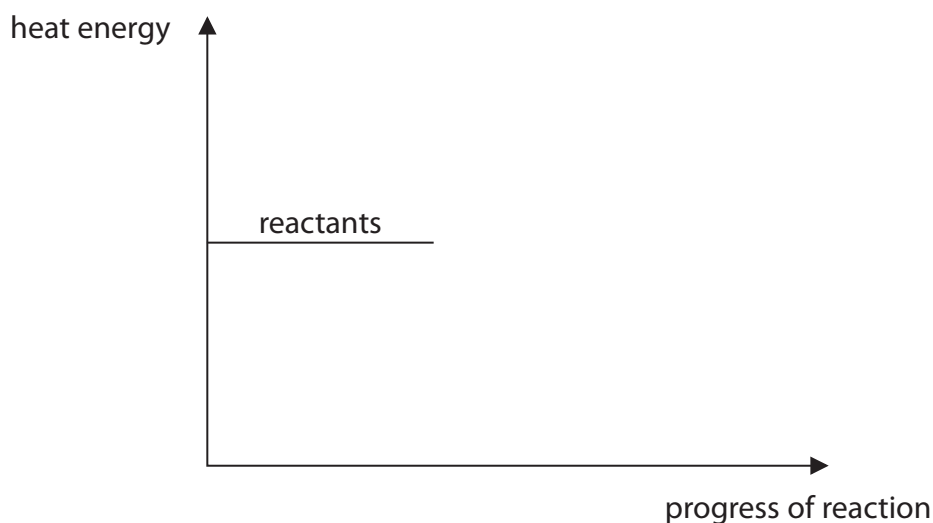
- (b) When zinc reacts with copper sulfate solution, copper and zinc sulfate solution are formed.



This reaction is exothermic.

Use this information to complete the diagram.

(2)



(c) Reactions are accompanied by heat changes.

The heat changes are the results of bonds being broken and bonds being formed.

Which row of the table shows the heat energy changes that occur when bonds are broken and when bonds are formed?

Put a cross (☒) in the box next to your answer.

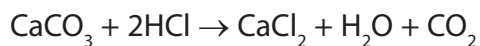
(1)

	<b>bonds broken</b>	<b>bonds formed</b>
<input type="checkbox"/> <b>A</b>	heat energy is released	heat energy is released
<input type="checkbox"/> <b>B</b>	heat energy is required	heat energy is required
<input type="checkbox"/> <b>C</b>	heat energy is released	heat energy is required
<input type="checkbox"/> <b>D</b>	heat energy is required	heat energy is released



3 (a) Marble chips react with hydrochloric acid to produce carbon dioxide.

The equation for the reaction is



Which one of these changes would **decrease** the rate of this reaction?

Put a cross (☒) in the box next to your answer.

(1)

- A** use hydrochloric acid which is more dilute
- B** use smaller sized marble chips
- C** use marble chips which have a larger surface area
- D** use a larger volume of the hydrochloric acid

(b) Explain why increasing the temperature of a reaction increases the rate of the reaction.

(2)

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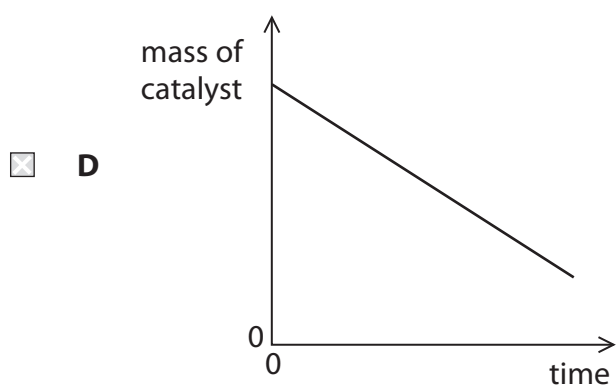
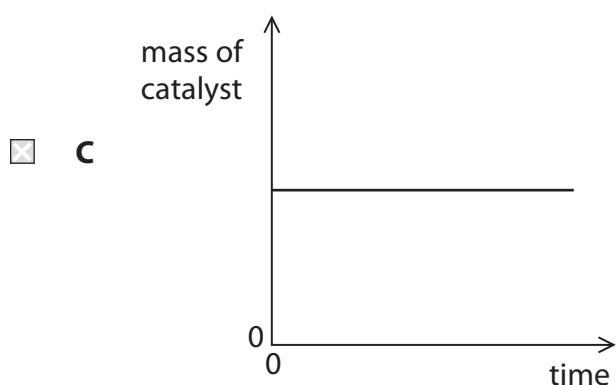
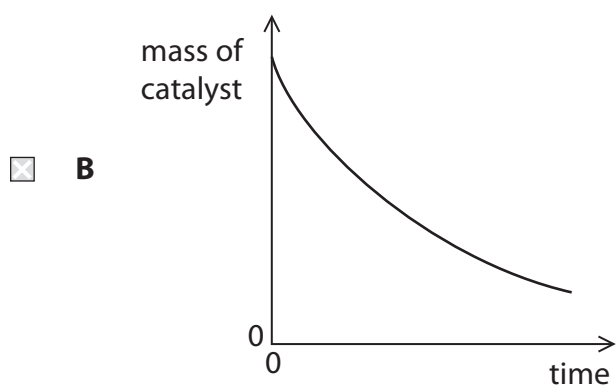
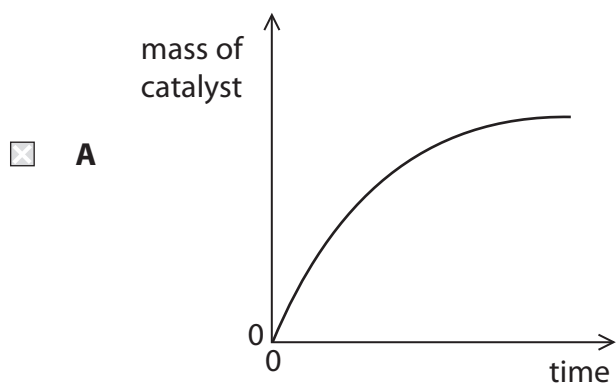


(c) (i) The rate of decomposition of hydrogen peroxide can be increased by adding a catalyst.

Which of these graphs shows the mass of the catalyst during the reaction?

Put a cross (☒) in the box next to your answer.

(1)



(ii) The decomposition of hydrogen peroxide,  $\text{H}_2\text{O}_2$ , produces oxygen and water.

Give the balanced equation for this reaction.

(2)

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(d) Explain, in terms of the energy involved in the breaking of bonds and in the making of bonds, why some reactions are exothermic.

(2)

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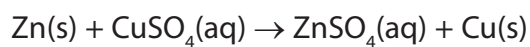
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**(Total for Question 3 = 8 marks)**

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- (b) Zinc is reacted with copper sulfate solution.  
The equation for the reaction is



- (i) What type of reaction is this?

(1)

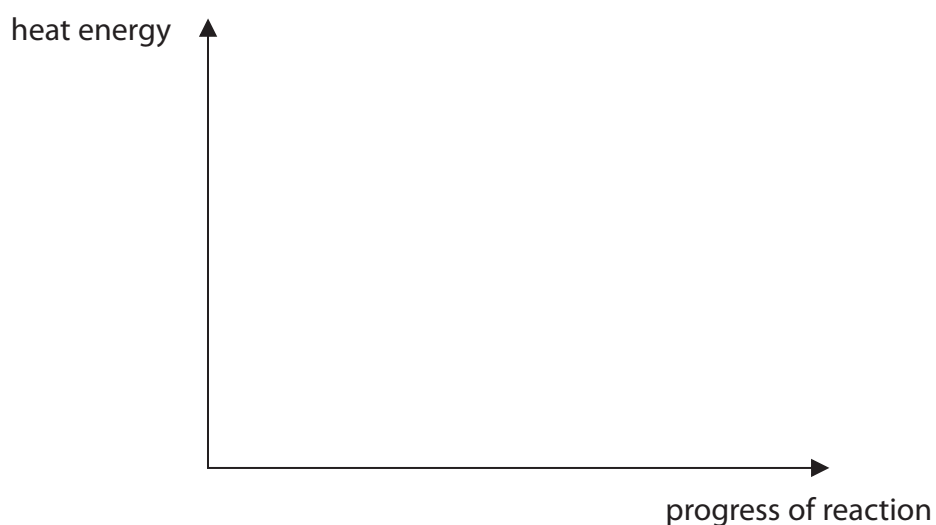
Put a cross (☒) in the box next to your answer.

- A decomposition
- B displacement
- C dissolving
- D neutralisation

- (ii) This reaction is exothermic.

On the diagram below draw labelled lines to show the relative energies of the reactants and products in this reaction.

(2)



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**(Total for Question 4 = 11 marks)**