

Mark Scheme (Results)

January 2016

International GCSE

Chemistry (4CH0) Paper 1C

Science Double Award (4SC0) Paper 1C

Pearson Edexcel Certificate in

Chemistry (KCH0) Paper 1C

Science (Double Award) (KSC0) Paper 1C

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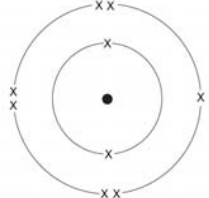
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General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.

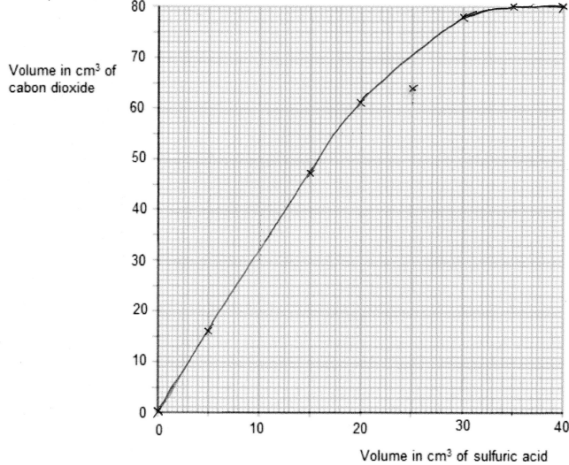
Question number	Answer	Notes	Marks
1 (a) (i)	A (Ag)		1
	(ii) D (Zr)		1
1 (b) (i)	3	'energy level' for 'shell' ignore references to inner shells ignore 'it has a valency of 3' 'energy levels' for 'shells' accept 'it has three shells'	1
(ii)	(The atom has) three <u>electrons</u> in its outer / valence shell		1
(iii)	3		1
(iv)	(The atom has) electrons in three shells / three shells are occupied (with electrons)		1
(v)	aluminium / Al		1
(c)		accept any symbol for electrons, eg dots, the letter 'e'	1

Question number	Answer	Notes	Marks												
2 (a)	C (halogens)		1												
(b) (i)	<p>M1 <u>atoms</u> of the same element</p> <p>M2 with different masses</p>	<p>accept '<u>atoms</u> with the same atomic number' / '<u>atoms</u> with the same number of protons'</p> <p>accept 'different mass numbers' / 'different numbers of neutrons'</p> <p>ignore references to electrons unless incorrect</p>	1 1												
(ii)	<table border="1" data-bbox="439 754 1113 1026"> <thead> <tr> <th data-bbox="439 754 595 887">Isotope</th> <th data-bbox="595 754 752 887">Number of protons</th> <th data-bbox="752 754 931 887">Number of neutrons</th> <th data-bbox="931 754 1113 887">Number of electrons</th> </tr> </thead> <tbody> <tr> <td data-bbox="439 887 595 954">${}_{35}^{79}\text{Br}$</td> <td data-bbox="595 887 752 954">35</td> <td data-bbox="752 887 931 954">44</td> <td data-bbox="931 887 1113 954">35</td> </tr> <tr> <td data-bbox="439 954 595 1026">${}_{35}^{81}\text{Br}$</td> <td data-bbox="595 954 752 1026">35</td> <td data-bbox="752 954 931 1026">46</td> <td data-bbox="931 954 1113 1026">35</td> </tr> </tbody> </table>	Isotope	Number of protons	Number of neutrons	Number of electrons	${}_{35}^{79}\text{Br}$	35	44	35	${}_{35}^{81}\text{Br}$	35	46	35		3
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${}_{35}^{79}\text{Br}$	35	44	35												
${}_{35}^{81}\text{Br}$	35	46	35												

	M1 first column correct M2 second column correct M3 third column correct		
(c)	ethane – no change (in colour)	accept '(stays) orange' ignore 'no reaction' /'nothing happens'	1
	ethene – (orange to) colourless / decolourises	ignore 'discolours' ignore starting colour of bromine	1

Question number	Answer	Notes	Marks
3 (a)	nitrogen / N ₂	accept N	1
(b)	oxygen AND water	accept steam	1
(c)	incomplete combustion (of the octane / fuel)	accept '(burns in a) limited supply / shortage of oxygen/air' reject 'no oxygen'	1
(d) (i)	$N_2 + 2O_2 \rightarrow 2NO_2$	accept halves and multiples accept as two correct equations via NO	1
(ii)	(It produces) acid rain OR (it causes) breathing problems / asthma	accept 'photochemical smog' ignore refs to greenhouse gas / global warming / climate change ignore refs to pollution	1

Question number	Answer	Notes	Marks
4 (a)	water	accept H ₂ O accept water vapour if both name and formula given mark name only	1
(b)	carbon dioxide	accept CO ₂ if both name and formula given mark name only	1
(c)	M1 (the copper / it) reacts with oxygen / oxidises M2 to form copper(II) oxide (which is black)	accept 'combines with/joins with/burns in oxygen' ignore 'air' accept 'copper oxide' reject 'copper(I) oxide'	2

Question number	Answer	Notes	Marks
5 (a)	 <p data-bbox="398 831 965 895">M1 & M2 all points correctly plotted to nearest gridline</p> <p data-bbox="398 975 965 1038">M3 suitable curve of best fit, from the origin</p>	<p data-bbox="1048 831 1794 967">deduct one mark for each incorrectly plotted point do not penalise missing (0, 0) if points are not visible, but graph goes through that point, then do not penalise</p>	3

Question number	Answer	Notes	Marks
(b) (i)	25 (cm ³)	accept anomalous point based on graph drawn	1
(ii)	<p>M1 the volumes (of gas) are the same</p> <p>M2 therefore the reaction has finished / <u>all</u> of the solid/MgCO₃ has reacted / the solid/MgCO₃ has been used up</p>	<p>accept 'no more gas is being produced/collected (after 35 cm³)'</p> <p>reject 'all of the reactants have reacted'</p> <p>reject 'all of the acid has reacted'</p> <p>ignore refs to MgCO₃ dissolving</p> <p>accept refs to MgCO₃ being limiting reagent</p>	2
(iii)	value correctly read to nearest gridline from candidate's graph		1
(iv)	value correctly read to nearest gridline from candidate's graph		1

Question number	Answer	Notes	Marks
6 (a)	(i) $2\text{HgO} \rightarrow 2\text{Hg} + \text{O}_2$	accept halves and multiples	1
	(ii) redox	accept '(thermal) decomposition' ignore 'oxidation' allow 'reduction'	1
(b)	(i) (tap / dropping / separating) funnel	reject 'filter / thistle funnel'	1
	(ii) (the gas / it) contains air (from the conical flask)	accept 'contains impurities' or ref to possible named impurity eg nitrogen reject 'water vapour' allow 'contains less <u>oxygen</u> '	1
(c)	<p>M1 perform reaction with and without catalyst</p> <p>M2 keep remaining variables (eg concentration or volume of hydrogen peroxide / temperature) the same</p> <p>M3 measure time (to fill the gas jar with oxygen)</p> <p>M4 <u>oxygen produced</u> more quickly/at a faster rate/in a shorter time (in experiment) with catalyst</p> <p>OR</p> <p>M1 weigh a sample of manganese(IV) oxide</p>	<p>accept:</p> <p>M1 perform reaction with and without catalyst</p> <p>M2 <u>oxygen produced</u> more quickly/at a faster rate/in a shorter time (in experiment) with catalyst</p> <p>M3 weigh a sample of manganese(IV) oxide (before putting it into the conical flask)</p> <p>M4 the mass at the end of the reaction should be the same as at the start</p>	4

	<p>(before putting it into the conical flask)</p> <p>M2 filter (to remove the solid)</p> <p>M3 dry the solid (and re-weigh it)</p> <p>M4 the mass should be the same as before</p>		
(d) (i)	$\text{SO}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_3$	<p>accept $\text{SO}_2 + \text{H}_2\text{O} + \frac{1}{2}\text{O}_2 \rightarrow \text{H}_2\text{SO}_4$</p> <p>allow products shown as correct ions</p>	1
(ii)	<p>M1 (Universal Indicator turns) orange/yellow</p> <p>M2 (the solution/it) is acidic / contains hydrogen ions / contains H^+ ions</p>	<p>accept 'red'</p> <p>allow 'contains sulfurous / sulfuric acid'</p>	2

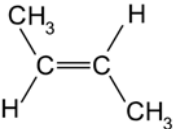
Question number	Answer	Notes	Marks
7 (a)	<p>M1 (Curve) A</p> <p>M2 faster reaction (at higher temperature)</p> <p>M3 therefore curve is steeper / curve levels off sooner</p>	<p>M2 and M3 dep on correct or missing M1 accept 'reaction takes less time'</p>	3
(b)	<p>M1 (Curve) C</p> <p>M2 only half the mass/amount of zinc used</p> <p>M3 therefore only half the volume / 20 cm³ of hydrogen produced</p>	<p>M2 and M3 dep on correct or missing M1 accept 'less zinc used, so less hydrogen produced' for 1 mark, if M2 and M3 not scored</p>	3

Question number	Answer	Notes	Marks
8 (a)	<p>(because) a precipitate was formed/a reaction took place each time Y was used</p> <p>OR</p> <p>no precipitate was formed/no reaction took place when X and Z were added together</p>	<p>accept 'it reacts with X and Z (to form a precipitate)'</p> <p>allow use of correct names for X, Y and Z</p>	1
(b)	<p>M1 X is (sodium) iodide and Z is (sodium) chloride</p> <p>M2 because X gives yellow precipitate or Z gives white precipitate</p> <p>OR</p> <p>M1 X is (sodium) iodide because it forms a yellow precipitate</p> <p>M2 therefore Z is (sodium) chloride</p> <p>OR</p> <p>M1 Z is (sodium) chloride because it forms a white precipitate</p> <p>M2 therefore X is (sodium) iodide</p>		2

(c)	M1 no change/no reaction with (sodium) chloride M2 colour change (to brown solution) with (sodium) iodide	accept 'orange' / 'orange-brown' accept 'grey/black <u>precipitate</u> ' reject incorrect colour change	2
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Question number	Answer	Notes	Marks
9 (a)	<p>M1 coke</p> <p>M2 limestone</p> <p>accept answers in either order</p>	<p>ignore 'carbon' / 'charcoal'</p> <p>ignore 'calcium carbonate'</p> <p>ignore formulae</p>	2
(b) (i)	<p>$\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$</p> <p>M1 all formulae correct</p> <p>M2 balanced</p>	<p>M2 dep on M1</p> <p>M2 dep on M1 or near miss eg Fe_2O_3 accept 'iron (III) ions / Fe^{3+} has gained electrons' accept 'oxidation number of iron decreases / oxidation number of iron changes from +3 to 0'</p>	2
(ii)	<p>M1 iron / Fe</p> <p>M2 (it has) lost oxygen</p>		2
(c) (i)	$\text{C} + \text{O}_2 \rightarrow \text{CO}_2$		1
(ii)	C (neutralisation)		1

(d)	(i)	M1 oxygen M2 water	accept 'air' accept 'moisture' / 'water vapour' ignore 'steam' accept answers in either order	2
	(ii)	prevents oxygen/water from coming into contact with the iron	accept refs to acting as a barrier	1
(e)	(i)	galvanising	ignore 'sacrificial protection'	1
	(ii)	M1 zinc is more reactive than iron / loses electrons more readily M2 (and therefore) corrodes in preference (to the iron)	ignore 'sacrificial protection' accept 'reacts (with oxygen/water) in preference (to the iron)' accept refs to zinc converting iron(II) ions to iron (atoms) reject 'zinc rusts' for M2	2
(f)	(i)	(aluminium/it) is too reactive / more reactive than carbon / <u>above</u> carbon in the reactivity series	accept 'carbon is less reactive than aluminium' accept 'the temperature required is too high' ignore refs to carbon monoxide	1
	(ii)	energy costs are too great / electricity is expensive	allow 'it is cheaper to use the blast furnace' ignore refs to iron being below carbon in the reactivity series	1

Question number	Answer	Notes	Marks
10 (a)	(the molecule) contains a (carbon to carbon) double bond	accept 'multiple bond' ignore refs to single bonds	1
(b) (i)	C_8H_{18} <u>and</u> C_2H_4	Ignore names of compounds	1
(ii)	M1 600-700°C		2
	M2 silica / alumina (catalyst)	accept 'aluminium oxide / silicon dioxide / aluminosilicate / zeolite' accept correct formulae	
(c) (i)	M1 (they have) the same <u>molecular</u> formula	allow 'both have same number of carbon and hydrogen (atoms as each other)' accept 'the atoms are arranged differently'	2
(ii)	M2 (but have) different structural formulae / displayed formulae / structures	accept  ignore bond angles accept fully displayed formula	1

10 (d) (i)	poly(propene) / polypropene (ii) $ \begin{array}{c} \text{CH}_3 \quad \text{H} \\ \quad \\ -\text{C} - \text{C}- \\ \quad \\ \text{H} \quad \text{H} \end{array} $ <p>M1 correct structure M2 extension bonds</p>	accept 'polypropylene' ignore brackets and 'n' M2 dep on M1 except award M2 if >1 repeat unit given	1 2
(e)	$ \begin{array}{c} \text{H} \quad \text{COOCH}_3 \\ \quad \\ \text{C} = \text{C} \\ \quad \\ \text{H} \quad \text{CN} \end{array} $	penalise incorrect use of upper / lower case letters and subscripts penalise bonds to incorrect atoms	1

Question number	Answer	Notes	Marks
11 (a)	<p>M1 chromate (ions) are negative</p> <p>M2 so they are attracted/move towards positive electrode/electrode B</p>	<p>accept 'anions'</p> <p>accept 'anode'</p>	2
(b) (i)	2 2 (1) (1)	accept halves and multiples	1
(ii)	B (HCl(aq))		1
(c) (i)	aq aq aq s	Do not accept words eg aqueous	1
(ii)	<p>M1 filter (off the precipitate)</p> <p>M2 wash (with distilled/deionised/pure water)</p> <p>M3 dry in a warm oven / leave to dry / dry with filter paper</p>	<p>allow 'decant'</p> <p>reject refs to crystallisation for M2 and M3</p> <p>allow 'heat it'</p>	3

Question number	Answer	Notes	Marks
12 (a) (i)	<p>M1 $0.53 \div 106$</p> <p>M2 $0.005(0)$ (mol)</p>	correct answer scores (2)	2
(ii)	<p>M1 $n(\text{CO}_2) = 0.005$ mol / answer to (a)(i)</p> <p>M2 $\text{vol}(\text{CO}_2) = (110 \div 0.005) = 22\,000$ (cm³)</p> <p>OR $110 \div$ M1 correctly evaluated</p>	correct answer scores (2)	2
(b)	<p>any two from:</p> <p>M1 the bung was not replaced quickly after the acid was added (so some carbon dioxide/gas escaped)</p> <p>M2 (some) carbon dioxide/gas dissolved in the water (in the trough or in the acid)</p> <p>M3 sodium carbonate is not pure</p>	<p>allow 'the bung was not on tightly/there was a leak around the bung (so some carbon dioxide/gas escaped)'</p> <p>allow 'reacted with the water'</p>	2

Question number	Answer	Notes	Marks
13 (a)	potassium / sodium / magnesium / zinc	accept K / Na / Mg / Zn if both name and symbol given, mark name only	1
(b)	M1 bubbles of gas produced rapidly/quickly M2 solid disappears quickly	accept any indication that the rate of evolution of bubbles and the disappearance of the solid is in between that of magnesium and zinc	2
(c) (i)	potassium hydroxide	accept KOH if both name and formula given, mark name only	1
(ii)	MgO		1
(d) (i)	carbon/C <u>and</u> it displaces/replaces zinc/Zn	reject 'displaces zinc oxide / displaces oxygen' accept 'it gains oxygen (from the zinc oxide) / it reduces zinc (oxide)'	1
(ii)	M1 carbon / C M2 it removes oxygen from the zinc (oxide) / causes zinc <u>ions</u> to gain electrons / gains oxygen / is oxidised	M2 dep on M1 reject 'displaces oxygen'	2

Question number	Answer	Notes	Marks
14 (a)	<p>M1 (goes darker because) more NO_2 is formed</p> <p>M2 as equilibrium/reaction shifts to left</p> <p>M3 because there are more moles/molecules (of gas) on the left hand side</p>	<p>allow 'moves backwards/in reverse direction'</p> <p>accept 'fewer moles/molecules on the right hand side'</p> <p>ignore references to Le Chatelier's principle</p>	3
(b) (i)	<p>M1 the equilibrium/reaction has shifted to the right / more N_2O_4 has been formed</p> <p>M2 a decrease in temperature shifts the equilibrium in the exothermic direction</p>	<p>accept 'therefore the (forward) reaction is exothermic' for M2 if M1 has been awarded</p>	2
(ii)	<p>(yes: because) bond making is exothermic/releases (thermal/heat) energy</p>		1

Question number	Answer	Notes	Marks
15 (a)	$3\text{Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2$ <p>M1 formula for magnesium nitride correct M2 rest of equation correct</p>	M2 dep on M1	2
(b) (i)	<p>M1 (damp) red litmus (paper) M2 turns blue</p> <p>OR</p> <p>M1 mix with hydrogen chloride/HCl M2 white solid/smoke forms</p>	<p>reject 'blue litmus' for both M1 and M2</p> <p>accept any suitable indicator with correct colour change, eg phenolphthalein turns red/pink</p> <p>reject 'hydrochloric acid' / 'HCl(aq)' but accept 'fumes from conc. hydrochloric acid'</p> <p>ignore 'fumes'</p>	2

(b) (ii)	<p>M1 M_r of lithium nitride = 35</p> <p>M2 $(1.40 \div 35 =) 0.04(0)$ (mol)</p>	correct answer scores (2)	2
(iii)	<p>M2 from (b)(ii) $\times 3 / 0.04(0) \times 3 = 0.12$ (mol)</p>		1
(iv)	<p>Using answer to b(iii)</p> <p>M1 answer to (b)(iii) $\div 2 / 0.12 \div 2 = 0.06(0)$ (mol)</p> <p>M2 answer to M1 $\div 0.500 / 0.06(0) \div 0.500$</p> <p>M3 $0.12 \text{ dm}^3 / 120 \text{ cm}^3$</p> <p>Using answer to b(ii)</p> <p>M1 answer to (b)(ii) $\div 2 / 0.04(0) \div 2 = 0.02(0)$ (mol)</p> <p>M2 answer to M1 $\div 0.500 / 0.02(0) \div 0.500$</p> <p>M3 $0.04 \text{ dm}^3 / 40 \text{ cm}^3$</p>		3

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